



A nanoscale p–n junction photoelectrode consisting of an NiO_x layer on a TiO₂/CdS nanorod core-shell structure for highly efficient solar water splitting

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ABSTRACT

The TiO₂/CdS system has attracted great attention in solar water-splitting applications owing to its desirable electronic and optical properties. With the aim of enhancing its photoelectrochemical water splitting efficiency, an efficient strategy is proposed via nanostructuring and linking it in a p–n junction configuration with NiO_x. The deposition of TiO₂ nanorods (NRs) and CdS was achieved using a hydrothermal synthesis route in the sequence, after which NiO_x was deposited via RF magnetron sputtering. Characterisation revealed the uniform deposition of CdS onto the TiO₂ NRs, forming a core-shell morphology, and the deposition of NiO_x on top of the TiO₂-NR/CdS resulted in a nanostructured p–n junction. X-ray photoelectron spectroscopy was used to resolve the valence band edge, and impedance studies confirmed the formation of a p–n junction; accordingly, the probable band edge positions of the photoelectrode were identified. The optimised TiO₂-NR/CdS-NiO_x p–n junction electrode exhibited a remarkable photocurrent of ~30 mA cm⁻² (at 1 V vs. Ag/AgCl) under AM 1.5 G simulated sunlight and an incident photon-to-current efficiency of ~97% at 500 nm. Furthermore, during illumination, the production of H₂ gas occurred with a faradaic efficiency of 95%. The results of the study demonstrate the advantage of utilizing the TiO₂-NR/CdS-NiO_x system in a p–n junction configuration to greatly enhance the charge generation, separation and suppression of the charge recombination, which boosts its photoelectrochemical water-splitting performance.

1. Introduction

Hydrogen is recognised as a promising clean energy carrier of the future that could replace depleting fossil fuels and decrease the serious environmental damage associated with them [1]. Moreover, sustainable hydrogen production technology using renewable sources could facilitate its commercial acceptance as a fuel in the future. The photoelectrochemical (PEC) water-splitting process using semiconductors is a promising approach to the carbon-free production of hydrogen and a sustainable solution for the future energy supply [2]. After the successful demonstration of photo-assisted water splitting using TiO₂ by Fujishima and Honda [3], tremendous effort has been devoted for introducing and investigating various semiconductor materials for efficient PEC water splitting. However, hydrogen production via efficient PEC water splitting remains a scientific challenge.

Among the metal oxide semiconductors, TiO₂ is the most widely studied material for PEC hydrogen production (via both PEC and

photocatalytic water splitting) owing to its well-matched conduction and valence band (VB) edge potential (with respect to water oxidation and reduction potential), low cost, photo-corrosion resistance and high stability in aqueous conditions [4,5]. Despite these advantages, the PEC activity of TiO₂ face major limitations due to its wide band gap (3.2 eV), that allows the absorption of only ultraviolet region of incident sunlight and incurs poor conductivity of photo-generated charge carriers, which hampers its reaction rate [6]. Hence, the solar water-splitting efficiency of TiO₂ tends to decrease when employed as a single-junction photoanode in a PEC system. Because of its wide band gap, TiO₂ is generally sensitised with low band gap semiconductors such as dyes, CdS etc to extend its light absorption ability [7,8]. In fact, TiO₂ is the most widely utilized support material for the sensitization of variety of narrow band gap semiconductors in numerous energy related applications including, dye, quantum dot and perovskite solar cell. Importantly, as a support material, TiO₂ provides an effective pathway for the transport of photogenerated electrons from the sensitizer towards the conducting

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substrate for faster and improved charge separation. Furthermore, TiO_2 can be fabricated in variety of nanostructure thin films having specific advantageous for the efficient PEC water splitting. Most importantly, the 1-D morphology of TiO_2 is known to provide direct electrical pathways for the photogenerated charges along the nanorod arrays, while providing efficient charge separation across the diameter.

CdS is one of the most important materials utilised in photo-catalytical applications due to its relatively low work function, high carrier mobility and significant absorption of visible light with a band gap of ~ 2.3 eV [9,10]. With such advantageous, CdS has the potential to act as a key ingredient in PEC system as an efficient light harvesting material to produce high photocurrents and H_2 . However, CdS face limitations such as the rapid recombination of photogenerated electron-hole pairs and photo-corrosion that strongly affects its photocatalytic activity [11,12]. Therefore, resolving the photo-corrosion and increasing the charge separation of CdS could allow its effective utilization as a suitable photoanode material in PEC system for efficient H_2 production. This can be achieved by coating a thin layer of stable material on CdS that could avoid the photooxidation of CdS by rapid extraction of photogenerated holes and by the refraining from direct contact with the electrolyte. However, like TiO_2 the solar water-splitting efficiency of CdS tends to decrease when used as a single-junction photoanode in the PEC system.

The limitations of both TiO_2 and CdS can be resolved by combining them to form a heterojunction photoelectrode. By the sensitisation of CdS on TiO_2 , the overall optical absorption can be extended to the visible spectrum up to 600 nm (including the UV region) [7]. Furthermore, the TiO_2 /CdS heterojunction photoelectrode possesses a higher specific surface area and more active reaction sites compared to individual single-junction electrodes [13]. Importantly, the conduction band (CB) edge of CdS is situated more negative than that of TiO_2 , that allows efficient transfer of the photo-generated electrons from CdS to the CB of TiO_2 during the illumination, whereas the holes remain in the CdS VB as a result of the potential gradient at the TiO_2 /CdS heterojunction interface [8,13]. This mechanism suppresses the recombination between photo-generated electrons and holes through efficient charge separation and improves the overall photocatalytic activity. Thus, a TiO_2 /CdS heterojunction electrode performs significantly more efficient than the individual TiO_2 and CdS photoelectrodes [13,14]. However, the efficiency of TiO_2 /CdS heterojunction can be further improved by providing a suitable nanostructure for direct charge transfer and surface coating for efficient charge separation and stability [13,15,16].

In PEC water splitting system, various strategies can be implanted to improve the efficiency of photoelectrodes. For example, the formation of a p–n junction photoelectrode is considered an important strategy to achieve increased electron–hole lifetime, efficient electron–hole separation and the suppression of electron–hole recombination. Because the formation of a p–n junction produces a space–charge region at the electrode junction due to the diffusion of charges [17–19]. The diffusion of electrons and holes towards opposite directions produces an internal electric field at the heterojunction region. During the illumination and with the applied bias potential, the produced internal electrical field in a p–n junction can significantly promote the efficient charge separation and transfer at the interface to produce a high quantum yield [20–23]. Hence, the PEC water-splitting efficiency of n-type photoelectrodes can be increased by combining with a p-type material to form nanostructured p–n junction electrodes. Most importantly, such a coating will ensure the stability against photo-corrosion and control the surface energetics and kinetics.

NiO is a p-type semiconductor material with a wide band gap of 3.6–4.0 eV and possesses a large work function of 5.0 eV [24]. Recent studies show that NiO can be efficiently utilised as a hole transfer layer [25], an interfacial layer in a solar cell [26], a chemically stable coating, a highly active electrocatalyst for the water oxidation [27], a layer to control the surface energetics and kinetics of photoanodes [28]

and as a photocathode [29] in PEC water-splitting devices. Importantly, the effective hole extraction, high hole mobility and stability enables its efficient utilisation as a hole transfer layer and surface coating. Hence, an optimised layer of NiO on photoanodes can effectively suppress the charge recombination to achieve efficient charge separation during PEC water splitting [20,27,28,30]. Furthermore, when NiO is combined with variety of materials in a heterojunction system, the wide band gap of NiO does not restrict/obstruct the light harvesting properties of other narrow band gap semiconductor materials. Hence, NiO can also be utilized as surface coating layer for narrow band gap semiconductors for protection against the photo-corrosion. Therefore, inspired by NiO's properties and with the aim of improving the PEC activity in the TiO_2 -NR/CdS system, we herein propose a p–n junction photoelectrode that contains n-type TiO_2 NRs with a CdS core-shell structure and p-type NiO_x . Moreover, a literature survey revealed that the fabrication and a PEC study of TiO_2 -NR/CdS- NiO_x p–n junction photoelectrode has not yet been reported.

Utilizing a hydrothermal method in sequence, TiO_2 NRs coated with a CdS shell structure on a fluorine-doped tin oxide (FTO) substrate was fabricated. A thin layer of NiO_x was deposited onto the TiO_2 -NR/CdS via RF magnetron sputtering. The NiO_x -sputtered condition that produced high PEC activity was chosen as the optimal deposition condition. The characterisation results reveal that NiO_x deposition mainly occurred on the top of the TiO_2 -NR/CdS shell structure. Compared to the TiO_2 -NR/CdS photoanode, the optimised TiO_2 -NR/CdS- NiO_x heterojunction photoelectrode exhibited a photocurrent density of ~ 30 mA cm^{-2} (at 1 V vs. Ag/AgCl) under AM 1.5 G illumination. Importantly, an incident photon-to-current efficiency (IPCE) of up to 97% at 500 nm and the hydrogen production with a faradaic efficiency of 95% was achieved using the TiO_2 -NR/CdS- NiO_x photoelectrode. The presence of p–n junction in TiO_2 -NR/CdS- NiO_x photoelectrode was confirmed via Mott-Schottky and I–V measurements. Furthermore, the band edge positions of TiO_2 , CdS and NiO_x before and after the heterojunction formation were investigated using X-ray photoelectron spectroscopy (XPS) VB edge analysis and electrochemical techniques.

2. Experimental

Aligned TiO_2 nanorods on an FTO substrate were synthesized by employing the previously reported hydrothermal method [31]. In a typical synthesis, 25 mL of HCl (37%, Sigma Aldrich) was mixed with 25 mL of deionised water and stirred for 30 min, after which 0.9 mL of titanium tetra-n-butoxide $[(\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{O})_4\text{Ti}]$, 97.0% Kanto chemical Co., INC, Japan] was added and the mixture was further stirred for 30 min to obtain a transparent synthesis solution. This was then transferred to the Teflon liner of a stainless-steel autoclave, filling 70% of its volume. Previously cleaned FTO substrates (2.5×3 cm) were then placed inside the Teflon liner with the conductive side facing downwards. The hydrothermal synthesis was carried out at 150 °C for 5.5 h in an electric oven, after which the autoclave was allowed to cool down to room temperature. Next, the TiO_2 -coated FTO substrates were taken out, rinsed thoroughly with deionised water and dried in ambient air. Finally, they were annealed at 450 °C for 2 h for conversion to the crystalline phase (with ramping at 5 °C/min).

The CdS core-shell structure on the TiO_2 NRs was formed by utilizing a previously reported method with slight modifications. Cadmium acetate dihydrate $[(\text{CH}_3\text{COO})_2\text{Cd} \cdot 2\text{H}_2\text{O}]$, 98%, Daejung chemicals, Korea] was used as a Cd source while thiourea ($\text{CH}_4\text{N}_2\text{S}$, 99%, Sigma Aldrich) was used as an S source. The aqueous hydrothermal synthesis solution was obtained by the addition 5 mM of $(\text{CH}_3\text{COO})_2\text{Cd} \cdot 2\text{H}_2\text{O}$ and 10 mM $\text{CH}_4\text{N}_2\text{S}$ in 50 mL deionised water. The synthesis solution was transferred to a Teflon-lined stainless-steel autoclave, filling 70% of its volume. The previously synthesised TiO_2 NR-coated FTO substrates were then placed inside the autoclave the TiO_2 NRs facing downwards. The hydrothermal synthesis was performed at 120 °C for 1.5 h in an electric oven, after which the autoclave was

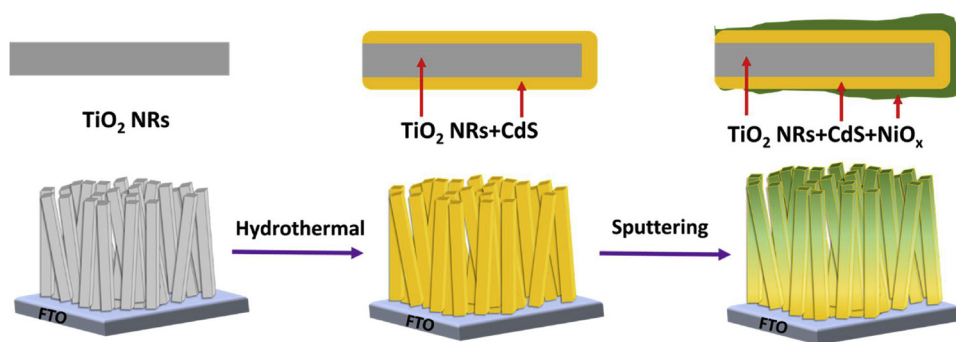


Fig. 1. The TiO₂-NR/CdS-NiO_x heterojunction photoelectrode fabrication scheme.

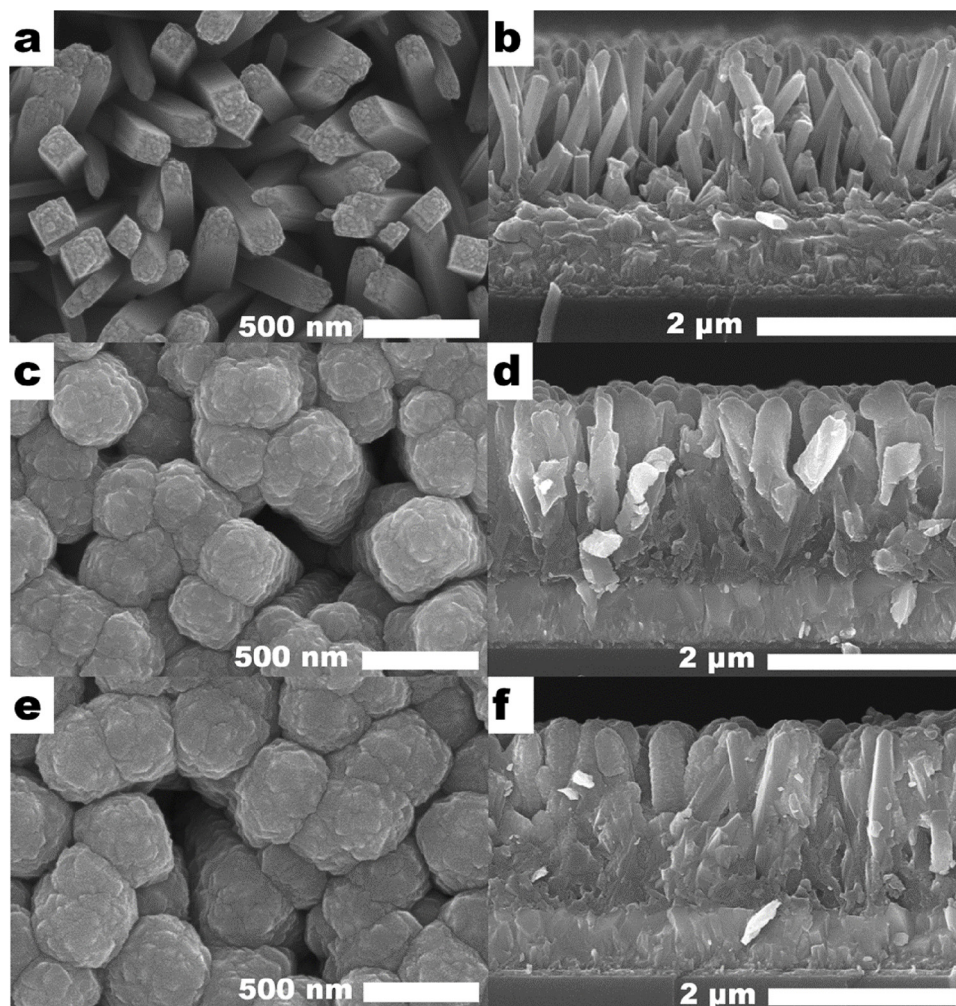


Fig. 2. Respective surface and cross-sectional SEM images of the (a and b) TiO₂ NR, (c and d) TiO₂-NR/CdS and (e and f) TiO₂-NR/CdS-NiO_x thin films.

allowed to cool and the CdS-coated TiO₂ NRs were washed and dried in air. A layer of NiO_x on the TiO₂/CdS was deposited via RF magnetron sputtering using a metallic Ni target (99.99% purity) in an Ar and O₂ gas mixture at a ratio of 20:7 (after optimisation). The deposition was carried out for 15 min (after optimisation) with an applied power of 100 W to the target. The substrates were masked with 3 M tape to avoid deposition in unwanted areas. The substrate (FTO-TiO₂-NR/CdS) was kept in rotation mode at 60 rpm without heating. Note that the optimal condition for NiO_x deposition was based on its PEC properties under varied deposition parameters such as gas ratio, flow rate, target power and deposition time. Fig. 1 shows the fabrication scheme for the TiO₂-NR/CdS-NiO_x p-n junction electrode.

The crystalline structures of the TiO₂/CdS-NiO_x thin films were characterised using a MiniFlex (Japan) desktop X-ray diffraction (XRD) instrument with Cu K α radiation at a diffraction angle between 10 and 90 2 θ . The surface and cross-sectional morphologies of the TiO₂/CdS-NiO_x thin films were analysed using a JEOL JSM-6700 F (Japan) and a Hitachi S4800 (Japan) scanning electron microscope (SEM). Transmission electron microscope (TEM) measurements were carried out using a JEOL JEM-2100 F (USA) instrument. UV-vis absorption spectra were recorded using a Varian Cary 5000 spectrophotometer (Australia) in integrated sphere scattering mode with a wavelength range of 200–2000 nm. An XPS analysis was carried out with a theta-probe angle-resolved X-ray photoelectron spectrometer (Thermo Fisher

Scientific Co., USA) using a monochromatic micro-focused Al K α (1486.6 eV) source; the peak energy in XPS was self-calibrated to the C 1 s and O 1 s reference peak states. An ABET Technologies Sun 2000 solar simulator (USA) calibrated to AM 1.5 G irradiance conditions was used as the light source. A Thermo Oriel 1000 W xenon arc lamp (Thermo Oriel Cornerstone 7400 1/8) was used to measure the IPCE. The amount of H₂ produced during the PEC activity was quantified using gas chromatography (YOUNG LIN - YL 6500 GC System, Korea) equipped with pulsed discharge ionization detector (PDD) containing capillary column. The gas collected in the sealed reaction cell was pumped into GC using helium carrier gas.

The PEC measurements of the TiO₂/CdS-NiO_x thin film photoanode were carried out in a quartz cell with Pt as the counter electrode and Ag/AgCl as the reference electrode in 0.5 M Na₂S and 0.5 M Na₂SO₄ electrolyte (separately) with a scan rate of 50 mV s⁻¹. The TiO₂/CdS-NiO_x photoanodes were masked with an O-ring during the PEC measurement, thereby exposing 0.125 cm² of the sample to the incident light source. An electrochemical impedance spectroscopic (EIS) analysis was conducted in the frequency range of 0.1 – 10⁵ Hz at 0 V vs. Ag/AgCl using an AC voltage perturbation of 10 mV in a 0.5 M Na₂S solution. Mott – Schottky measurements were also carried out in the same electrolyte solution in the absence of irradiation. The amount of hydrogen produced during electrolysis was measured using a calibrated pressure sensor. The applied potential was converted to RHE values using the following Nernst equation:

$$E_{\text{RHE}} = E_{\text{Ag/AgCl}} + (0.0591 \times \text{pH}) + E^{\circ}_{\text{Ag/AgCl}} \quad (1)$$

$$E^{\circ}_{\text{Ag/AgCl}} (3.0\text{M NaCl}) = 0.209\text{V at } 25^{\circ}\text{C}.$$

3. Results and discussion

Fig. 2a and b show the top and cross-sectional views of the TiO₂ NRs deposited on the FTO substrate, respectively. The TiO₂ NRs on FTO substrate were in tetragonal shape with square top facets aligned perpendicular to the substrate. The top surface of the TiO₂ NRs appears to contain many step edges, while the side wall appears smooth. The diameter and length of the TiO₂ NRs were found to be ~100 nm and ~1.8 μm , respectively. After the hydrothermal deposition of CdS onto the TiO₂ NRs, significant changes in morphology and dimension were observed. The top and cross-sectional SEM views shown in Fig. 2c and d confirm the complete coverage of CdS on the TiO₂ NRs forming a core-shell structure. After the CdS deposition, the overall dimension of the NRs appeared to have increased to 200–300 nm, whereas the length of the NRs increased to ~2 μm . Furthermore, the space between the NRs appeared to have narrowed due to CdS deposition. After the sputter deposition of NiO_x onto TiO₂-NR/CdS, no significant change was noticed in dimension and the length of NRs. Importantly, the sputter deposition conditions did not affect the morphology of TiO₂-NR/CdS, as evident in Fig. 2e and f. In these images (Fig. 2e and f), the NiO_x layers deposited on TiO₂-NR/CdS cannot be seen clearly due to the smaller particle size of NiO_x. After the magnification of surface SEM images, the change in surface roughness was noticed (Fig. S1) confirming the deposition of NiO_x nanoparticles on the top of TiO₂-NR/CdS. To confirm the nature of NiO_x deposition, it was sputtered onto the bare FTO and characterised. The Fig. S2 shows the surface and cross section SEM image of NiO_x deposited on the bare FTO substrate confirming the uniform deposition of NiO_x consisting of nanoparticles forming a thin-layer structure.

XRD patterns of the fabricated electrodes are shown in Fig. 3. All the XRD patterns contained peaks related to tetragonal SnO₂ (JCPDS No. 46-1088; indicated with black dots) that arise from the FTO substrate (green line). The XRD pattern of the synthesised TiO₂ NRs on FTO substrate (black line) was ascribed to the rutile phase of tetragonal TiO₂ (JCPDS No.88-1175) [31]. The appearance of XRD peaks related to

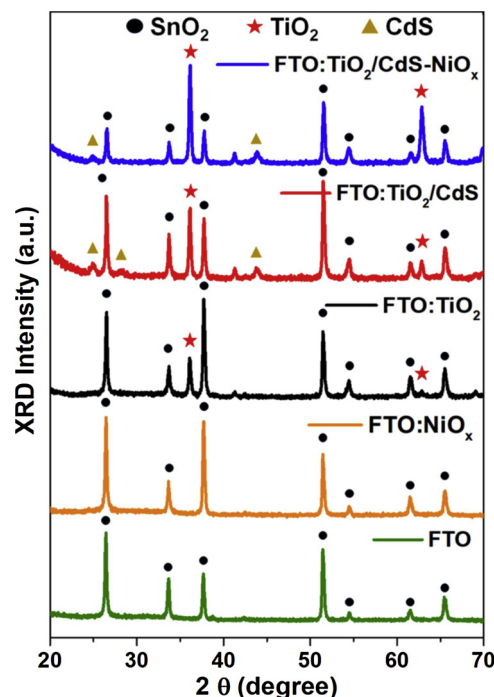


Fig. 3. XRD pattern of the FTO substrate and NiO_x, TiO₂ NR, TiO₂-NR/CdS and TiO₂-NR/CdS-NiO_x thin films on FTO substrates. The black dots indicate the XRD peaks that arise from SnO₂ in the FTO substrate, whereas the XRD peaks indicated by stars and triangles indicate the presence of rutile TiO₂ and hexagonal CdS, respectively.

(101) and (002) planes and the absence of other peaks indicate the preferential growth of TiO₂ NRs and its orientation with respect to the substrate [31]. In the XRD pattern of the TiO₂-NR/CdS (the red line in Fig. 3), the diffraction peaks arising from CdS can be clearly seen and were assigned to the hexagonal phase of wurtzite CdS (JCPDS card no. 75-1545) [32]. The observed characteristic diffraction peaks at 2 θ angles of ~24.95, ~26.51, ~28.17 and ~43.80 were respectively assigned to the (100), (002), (101) and (110) planes of the hexagonal CdS phase, respectively. However, the XRD pattern of the FTO-TiO₂-NR/CdS-NiO_x (blue line in Fig. 3) and FTO-NiO_x (orange line in Fig. 3) samples did not show any obvious XRD peaks of the NiO_x phase, suggesting that the NiO_x layer was very thin.

TEM was employed for the detailed microscopic characterisation and analysis of the TiO₂-NR/CdS-NiO_x junction. The low-resolution TEM image of the TiO₂ NRs shown in Fig. 4a reveals that the TiO₂ NR walls appeared smooth with an NR diameter of ~100 nm. The high-resolution TEM image (Fig. 4b) that shows clear lattice fringes confirms the high crystallinity of the TiO₂ NRs. From these, the interplanar d spacing value was calculated as 0.324 nm assigned to the (110) plane of the rutile TiO₂ phase [31]. The hydrothermal deposition of CdS on the TiO₂ NRs was found to have increased the overall NR diameter to ~275 \pm 25 nm with a corresponding change in the surface smoothness (Fig. 4c). The overall increase in the NR diameter indicates that the thickness of CdS was around 75–100 nm. Importantly, the deposition of CdS was uniform throughout the TiO₂ NR surface, thus confirming the formation of a core-shell structure. The interplanar d spacing value of 0.35 nm (Fig. 4d) obtained from the high-resolution TEM image was indexed to the (100) plane of the hexagonal CdS. Furthermore, the high-resolution TEM images in Fig. 4e and f clearly indicate the interface between the TiO₂ and CdS without any barrier, which would be beneficial for better charge transfer at the interface. Hence, it can be concluded that the proposed procedure yielded uniformly distributed and highly crystalline CdS on the TiO₂ NRs. The deposition of NiO_x onto TiO₂-NR/CdS showed no significant change in the NR diameter or

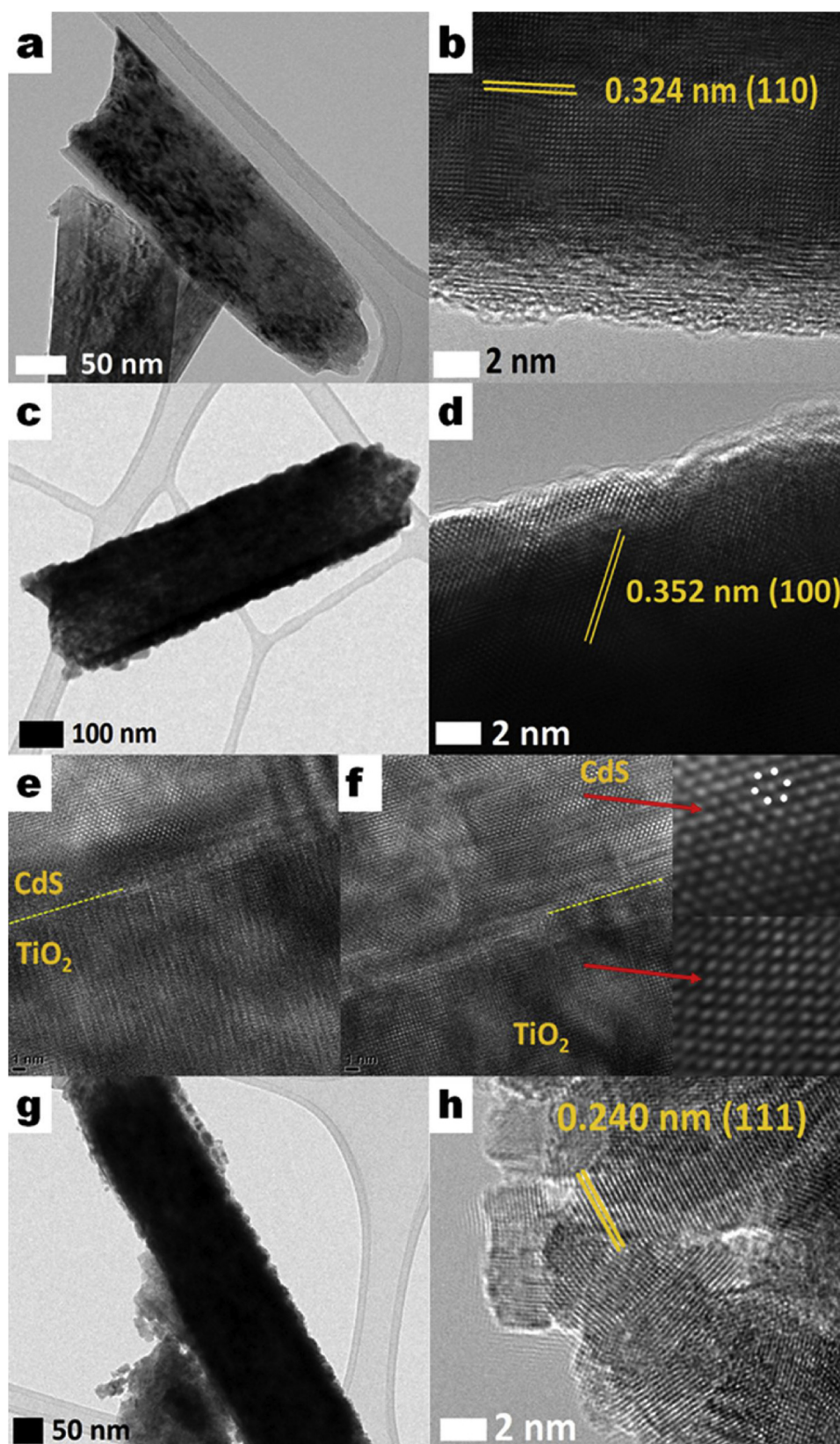


Fig. 4. Respective low- and high-resolution TEM images with d-spacing values of the (a and b) TiO₂ NRs, (c and d) TiO₂-NR/CdS and (g and h) TiO₂-NR/CdS-NiO_x; (e and f) high-resolution TEM images showing the interface between TiO₂ and CdS.

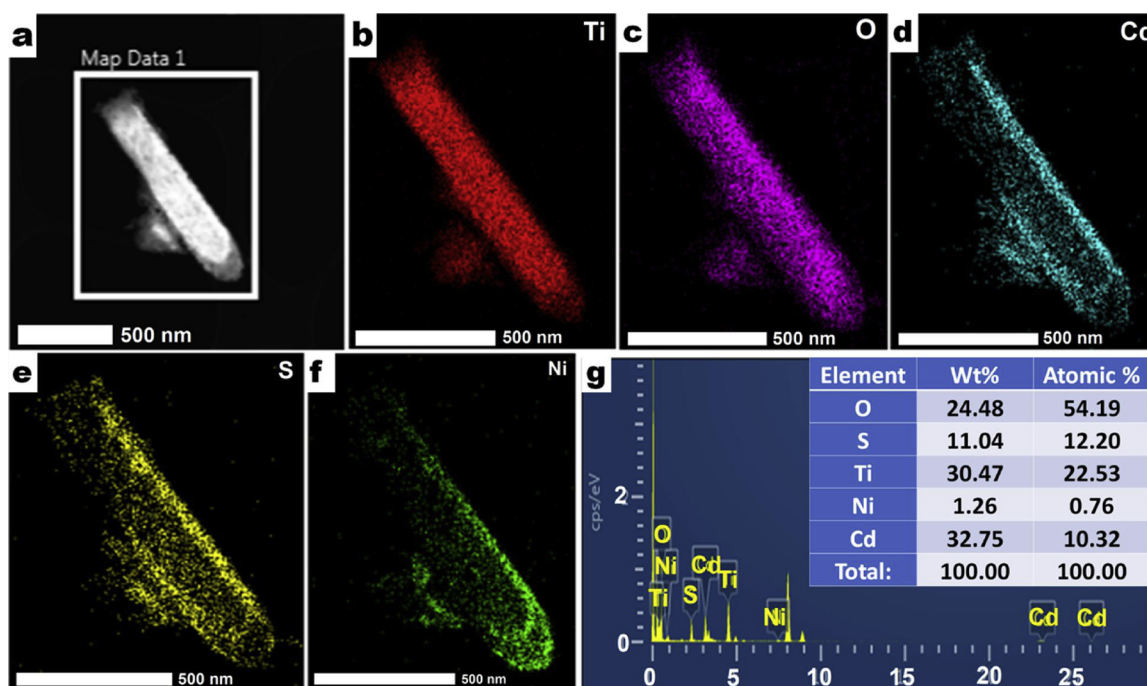


Fig. 5. (a) STEM image of $\text{TiO}_2\text{-NR/CdS-NiO}_x$ showing the selected area for elemental EDS mapping. Elemental mapping images of (b) Ti, (c) O, (d) Cd, (e) S and (f) Ni. (g) EDS spectrum of $\text{TiO}_2\text{-NR/CdS-NiO}_x$ in the selected area. The inset table shows the elemental analysis data.

surface roughness (Fig. 4g) due to the low amount of deposited NiO_x . The TEM analysis at various locations on the NRs revealed that noticeable amounts of NiO_x were deposited on the top compared to the surface or bottom of the $\text{TiO}_2\text{-NR/CdS}$. The interplanar d-spacing value of 0.24 nm was obtained from the lattice fringes (Fig. 4h), which was assigned to the (111) plane of the cubic phase of NiO_x [33]. The TEM images shown in Fig. S3 indicate that the particle size of NiO_x was between ~2–10 nm (with an average diameter of 5 nm). Due to the uneven deposition of NiO_x on the $\text{TiO}_2\text{-NR/CdS}$, the exact thickness of the NiO_x could not be determined using TEM. Furthermore, the interplanar d-spacing value of 0.20 nm observed in Fig. S3d was assigned to the (111) plane of metallic Ni [34], indicating the presence of highly crystalline NiO_x having multiple phases.

Energy-dispersive X-ray spectroscopy (EDS) mapping was employed to confirm the distribution and quantity of the individual elements present in $\text{TiO}_2\text{-NR/CdS-NiO}_x$ (Fig. 5 and S4). The analysis confirmed the presence of Ti, O, Cd, S and Ni without any impurities. The elemental mapping images in the selected area confirmed the uniform deposition of CdS on the TiO_2 NRs. Note that its thickness was dependent on the deposition time, as shown in Fig. S4. A careful observation of Fig. 5f confirmed that the deposition of NiO_x mainly occurred on the top of $\text{TiO}_2\text{-NR/CdS}$, with a very small amount deposited on the surface. The thickness of the NiO_x layer on the top of the $\text{TiO}_2\text{-NR/CdS}$ was ~70–100 nm.

The UV–Vis absorption spectra of the TiO_2 NRs, NiO_x , $\text{TiO}_2\text{-NR/CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ are shown in Fig. 6a. The bare TiO_2 NRs absorb only UV light with wavelengths < 410 nm (red line). Upon the deposition of CdS, the light absorption of TiO_2 was extended to the visible region up to 550 nm (blue line). However, the deposition of NiO_x on $\text{TiO}_2\text{-NR/CdS}$ caused slight red shift in the absorption onset (pink). As is evident (black line), the NiO_x thin films (NiO_x deposited on FTO) exhibited absorption onset at 350 nm with tailing in the visible region, confirming the presence of oxygen-deficient NiO_x and/or metallic Ni. The band gaps of the TiO_2 NR, NiO_x , $\text{TiO}_2\text{-NR/CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ photoelectrodes were determined by plotting $(\alpha h\nu)^2$ as a function of $h\nu$ as shown in Fig. 6b. In accordance with the absorption onsets, the band gaps of ~3.0, 3.5, 2.26 and 2.25 eV were obtained for the TiO_2 NR, NiO_x , $\text{TiO}_2\text{-NR/CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ photoelectrodes,

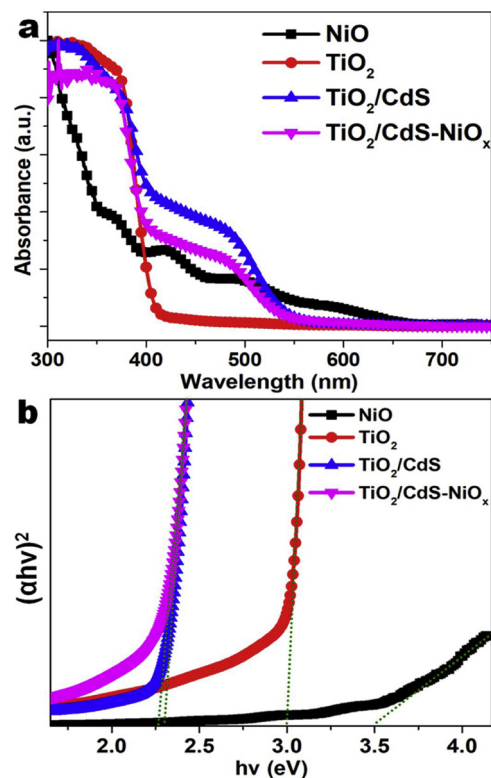


Fig. 6. (a) UV–vis absorbance spectra and (b) Tauc plots of NiO , TiO_2 NR, $\text{TiO}_2\text{-NR/CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ thin films on FTO substrates yielding band gaps of ~3.51, ~3.01, ~2.30 and ~2.26 eV, respectively.

respectively.

XPS was employed to determine the surface electronic state of the elements present in the photoelectrode, as shown in Fig. 7. The high-resolution XPS graph of Ti exhibited two peaks (Fig. 7a) ascribed to the Ti 2p_{3/2} and Ti 2p_{1/2} at a binding energy values of 458.7 and 464.4 eV,

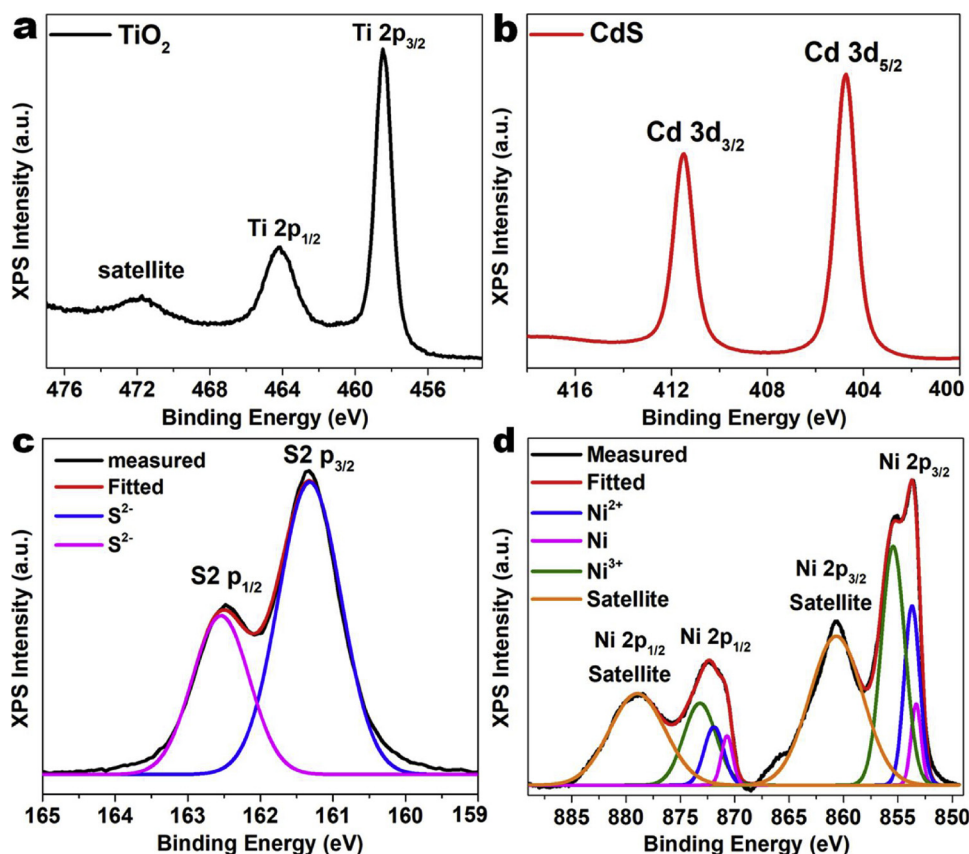


Fig. 7. High-resolution XPS spectra of (a) Ti 2p and (b) Cd 3d; deconvoluted high-resolution XPS spectra of (c) S 2p and (d) Ni 2p.

respectively which agree with the typical values of TiO₂ having the Ti⁴⁺ state [35]. Furthermore, the deconvolution of the Ti spectra did not produce any additional peaks, thereby confirming the presence of only the Ti⁴⁺ state in the TiO₂ NR. In addition, a high-energy shake up satellites at ~472.4 eV was noticed, indicating that the free energy of formation of TiO₂ is favoured over other titanium oxides as per the theoretical thermodynamics [36–38]. The high-resolution core level XP spectrum of Cd 3d exhibited two peaks (Fig. 7b) at binding energy values of ~405.00 and ~411.10 eV arising from Cd 3d_{3/2} and Cd 3d_{5/2}, respectively. This confirms the valency of Cd²⁺ in the CdS layer deposited on the TiO₂ NR [39]. Fig. 7c shows the high-resolution deconvoluted XP spectrum of S 2p. The S 2p spectrum exhibited two characteristic peaks arising from S 2p_{3/2} and S 2p_{1/2} at binding energy values of ~161.5 and ~162.7 eV, respectively, which are attributed to the sulphide moiety of CdS. Importantly, the deconvolution did not yield any peaks relevant to the disulphide moiety. The Ni exhibited characteristic XPS peaks arising from Ni 2p_{3/2} and Ni 2p_{1/2}. Interestingly, the deconvolution of Ni 2p XPS spectrum produced three peaks at 854.7, 856.0 and 852.9 eV (Fig. 7d) that respectively arose from Ni²⁺, Ni³⁺ and Ni⁰ states [40], confirming the presence of NiO, Ni₂O₃ and metallic Ni in the TiO₂-NR/CdS-NiO_x heterojunction photoelectrode.

The PEC properties of the electrodes were analysed using linear sweeping voltammetry (LSV) under AM 1.5 G simulated solar light irradiation (i.e. 100 mW cm⁻²) with the back-side illumination setup. By considering the stability of CdS, the PEC activities were tested in two different electrolytes: 0.5 M Na₂S and 0.5 M Na₂SO₄. The LSV voltammograms of the TiO₂ NR, TiO₂-NR/CdS and TiO₂-NR/CdS-NiO_x films in the presence and absence of illumination (dark) are shown in Fig. 8a. The TiO₂ NRs, TiO₂-NR/CdS and TiO₂-NR/CdS-NiO_x electrodes exhibited negligible/no photocurrent in the absence of illumination. The photocurrent density of TiO₂ NRs significantly increased after the deposition of CdS. The TiO₂ NRs and optimised TiO₂-NR/CdS (the CdS deposition time optimisation, see Fig. S5a) electrodes exhibited

respective photocurrent densities of 1.2 and 3.2 mA cm⁻². After the deposition of NiO_x on TiO₂-NR/CdS, a dramatic enhancement in the photocurrent was observed. Here, the amount of NiO_x deposited on TiO₂-NR/CdS can be controlled by varying the sputtering deposition time. The photocurrent of TiO₂-NR/CdS shows gradual increase with the increase in deposition time until 15 min. Above the 15 min. of deposition time tends to cause rapid decrease in photocurrent. Therefore, 15 min. of NiO_x deposition time was selected as optimum deposition condition. For the optimised NiO_x deposition for TiO₂-NR/CdS-NiO_x films (see Fig. S5b), a record photocurrent of ~30 mA cm⁻² was observed. Clearly, the photocurrent obtained for the TiO₂-NR/CdS-NiO_x electrode was significantly greater than that for the individual TiO₂ NRs and the TiO₂-NR/CdS electrodes. Interestingly, the deposition of NiO_x causes anodic shift in onset potential and comparatively lower photocurrents were observed at low bias voltages (-1.2 to -0.6 V) as noticed in Fig. 8a. It can be speculated that, the anodic shift in flat band potential might be caused by the surface states that effect the charge on the outer surface of the electrode which could lower the conduction band at the surface and hence shift onset potential anodically. Because, the synthesized NiO_x layer on TiO₂-NR/CdS consists of multiple phases that could result in a significant amount of surface states. Moreover, a large photocurrent observed at high bias could also be ascribed to the large drive force for the surface reaction (S²⁻ to S₂²⁻). Overall, the high photocurrent obtained for the optimised electrodes could have been due to the hole scavenging, charge transport properties of NiO_x forming p-n junction electrodes for efficient charge separation and migration along with the suppression of the charge recombination [41].

The IPCE values were obtained to study the photo-response of electrodes as a function of the incident light wavelength. Fig. 8b represents the IPCE plot of TiO₂ NRs, TiO₂-NR/CdS and TiO₂-NR/CdS-NiO_x electrodes measured with an applied bias of 0 V vs. Ag/AgCl in 0.5 Na₂S electrolyte as follows:

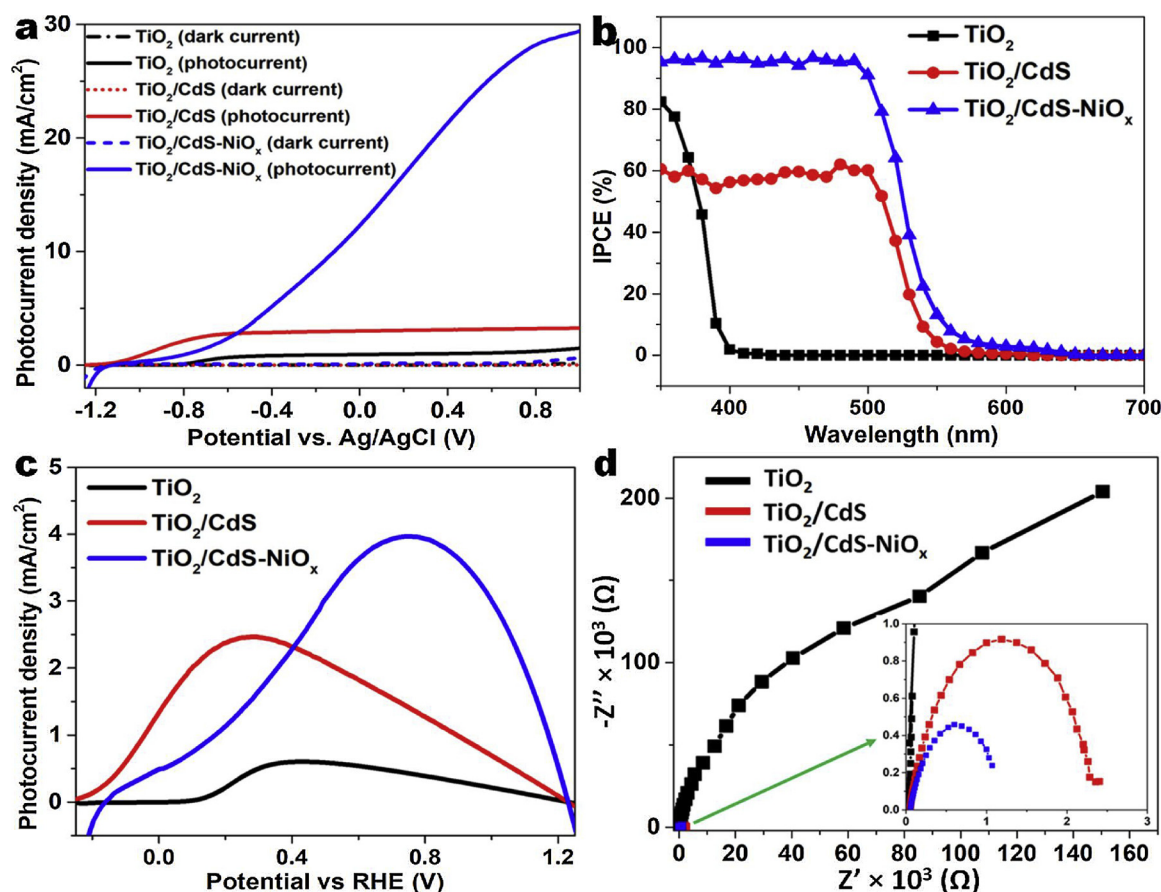


Fig. 8. (a) LSVs, (b) IPCEs, (c) photo-conversion efficiencies and (d) Nyquist plots of the TiO_2 NR, $\text{TiO}_2\text{-NR/CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ electrodes. IPCE and impedance were measured at an applied bias voltage of 0 V vs. Ag/AgCl in 0.5 M Na_2S solution. The full lines and the dotted lines in (a) represent the LSV graphs measured under the presence and absence of illumination, respectively for the corresponding photoelectrodes.

$$\text{IPCE} = [(1240 I) / (\lambda J_{\text{light}})] \times 100, \quad (2)$$

where I is the photocurrent density, λ is the incident light wavelength and J_{light} is the measured irradiance. In accordance with the absorbance results, the TiO_2 NR electrode did not produce IPCE beyond 400 nm. After the deposition of CdS on TiO_2 NR, the IPCE onset showed a significant red shift reaching maximum IPCE values of 60% at 500 nm. The subsequent deposition of NiO_x caused a dramatic enhancement in IPCE value to ~97% (at 500 nm) with a slight red shift in onset. These observations clearly confirm that the NiO_x deposition on $\text{TiO}_2\text{-NR/CdS}$ electrodes improved the visible-light absorption and enabled efficient charge separation and transportation, thereby producing a record enhancement in IPCE value in the visible spectrum. Fig. S6 shows the IPCE values and their corresponding integration plots for the TiO_2 NR, $\text{TiO}_2\text{-NR/CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ electrodes. The integration of the IPCE values in the wavelength range of 300–700 nm produced the expected photocurrent values that appeared at 0 V vs. Ag/AgCl, as shown in Fig. 8a, thus validating the measured photocurrent values under 1 sun illumination. Furthermore, the stabilities of the TiO_2 NR, $\text{TiO}_2\text{-NR/CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ electrodes were investigated by measuring the transient currents ($I-t$) with an applied bias of 0 V vs. Ag/AgCl under 1 sun illumination (Fig. S7). All of the electrodes showed a sudden increase in photocurrent value when illuminated after 1 min that stabilised over time. All the electrodes exhibited constant photocurrent values, even after 4 h of illumination, indicating their significant stability over the observed time range.

Fig. 8c represents the photo-conversion (i.e. photon-to-hydrogen) efficiencies of the electrodes calculated using the following equation [42]:

$$\eta (\%) = [I (E_{\text{rev}}^o - V)] / J_{\text{light}}, \quad (3)$$

where η is the photo-conversion efficiency, I is the photocurrent density (mA cm^{-2}), J_{light} is the incident light irradiance (mW cm^{-2}), E_{rev}^o is the standard reversible potential (1.23 V vs. RHE) and V is the applied bias potential vs. RHE. The $\text{TiO}_2\text{-NR}$ electrode exhibited the lowest photo-conversion efficiency of 0.72% at 0.41 V vs. RHE, while the $\text{TiO}_2\text{-NR/CdS}$ photoanode showed the efficiency of 2.5% at 0.30 V vs. RHE. Importantly, the deposition of NiO_x on $\text{TiO}_2\text{-NR/CdS}$ photoanode increased the photo-conversion efficiency to 4% at 0.78 V vs. RHE along with a positive shift in potential. Here, the observed positive shift in potential confirm the alteration in the electronic structures of TiO_2 NRs in $\text{TiO}_2\text{-NR/CdS}$ by the subsequent deposition of CdS and NiO_x , respectively [43]. Electrochemical impedance spectra (EIS) were acquired to evaluate the charge-transfer resistance across the electrode–electrolyte interface. The corresponding Nyquist plots of TiO_2 NRs, $\text{TiO}_2\text{-NR/CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ measured under illumination are presented in Fig. 8d. Compared to the TiO_2 NRs and the $\text{TiO}_2\text{-NR/CdS}$ electrode, $\text{TiO}_2\text{-NR/CdS-NiO}_x$ exhibited a smaller semicircle indicating that the deposition of NiO_x on $\text{TiO}_2\text{-NR/CdS}$ yielded lower charge transfer resistance across the electrode–electrolyte interface and thus better charge transfer across the interface, which could have enhanced the PEC activity.

The $\text{TiO}_2\text{-NR/CdS-NiO}_x$ p–n junction electrodes were tested for hydrogen evolution under simulated AM 1.5 G solar illumination in 0.5 M Na_2S electrolyte. Fig. 9a shows the H_2 gas production rates of the $\text{TiO}_2\text{-NR}$, $\text{TiO}_2\text{-NR/CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ photoelectrodes measured during 2 h of continuous irradiation. Compared to $\text{TiO}_2\text{-NR}$, the $\text{TiO}_2\text{-NR/CdS}$ show significantly higher hydrogen production rate due to the extended light absorption. Importantly, the rate of H_2 production

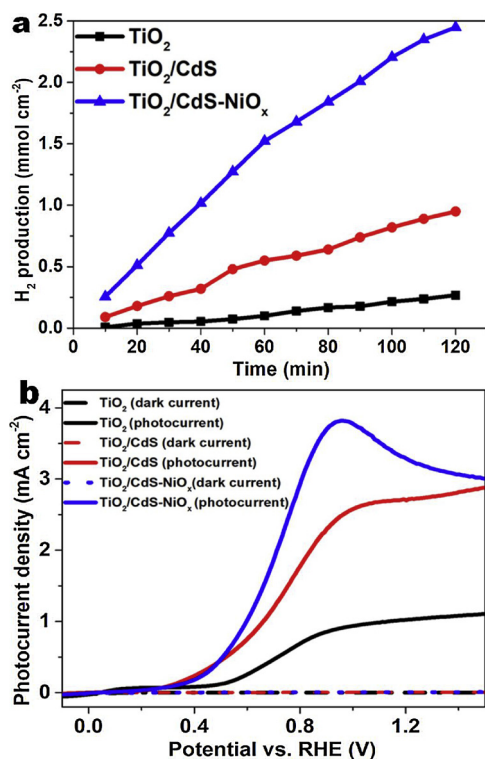


Fig. 9. (a) The production of hydrogen at the TiO₂-NR/CdS-NiO_x electrode (at 0 V vs. Ag/AgCl) under AM 1.5 irradiation in a 0.5 M Na₂S electrolyte. (b) LSVs of the TiO₂ NR, TiO₂-NR/CdS and TiO₂-NR/CdS-NiO_x electrodes measured in 0.5 M Na₂SO₄ under 1 sun illumination. The full lines and the dotted lines in (b) show the LSV graphs measured under illumination and dark condition, respectively for corresponding photoelectrodes.

increased dramatically after NiO_x deposition since the evolution of H₂ for TiO₂-NR/CdS and TiO₂-NR/CdS-NiO_x photoelectrodes were measured as 0.5 and 1.3 mmol h⁻¹ cm⁻², respectively. The obtained H₂ gas volumes and corresponding I-t curves were used to calculate the faradaic efficiencies of the TiO₂-NR/CdS and TiO₂-NR/CdS-NiO_x photoelectrode using the following equation:

$$\text{Faradaic efficiency} = n_{\text{H}_2}/(Q/2F), \quad (4)$$

where n_{H_2} is the total amount of hydrogen produced (mol), Q is the total amount of charge (C) and F is the Faraday constant. During the initial 1 h of hydrogen generation, the faradaic efficiencies of the TiO₂-NR/CdS and TiO₂-NR/CdS-NiO_x electrodes were calculated as 80 and 95%, respectively. However, after 2 h of irradiation, the faradaic efficiencies had noticeably decreased to 70 and 90%, respectively. Fig. 9b shows the LS voltammograms of the electrodes measured in 0.5 M Na₂SO₄ electrolyte. 0.5 M Na₂S acted as a sacrificial electron donor electrolyte via the oxidation of the sulphate moiety and protected the CdS from photo-corrosion, whereas 0.5 M Na₂SO₄ did not contain any sacrificial electron donors, and low comparative photocurrents were thus observed. Importantly, in the case of the 0.5 M Na₂SO₄ electrolyte, a significant increase in the photocurrent was observed after the NiO_x deposition on TiO₂-NR/CdS electrodes, indicating the advantages of the NiO_x layer on the CdS-based electrodes.

A detailed characterization of TiO₂NR/CdS-NiO_x electrode was performed after the PEC activity, to investigate its stability after the prolonged irradiation and H₂ production. The SEM images shown in Fig. S8 indicate that no significant change in the morphology was noticed after the PEC measurements. Fig. S9 shows the TEM characterization data of TiO₂NR/CdS-NiO_x electrode obtained after the PEC performance. The interplanar d spacing values of 0.321 and 0.359 nm observed in Fig. S9a were assigned to the (110) plane of the rutile TiO₂

and (100) plane of the hexagonal CdS phase, respectively. The interface between the TiO₂ and CdS was clearly seen in Fig. S9b without any barrier. The HR-TEM images shown in Fig. S9c confirm the particle sizes of NiO_x deposited on the top of TiO₂NR/CdS-NiO_x electrode after PEC performance. The interplanar d-spacing value of 0.239 nm and 0.200 nm observed in the Fig. S9d were ascribed (111) plane of the cubic and metallic NiO_x phase, respectively. Importantly, the TEM characterization indicate that the morphological and crystallographic properties of the TiO₂NR/CdS-NiO_x electrode remained stable and unchanged after the PEC performance. Furthermore, the elemental mapping and EDS was used to analyse the distribution and quantity of the individual elements in TiO₂NR/CdS-NiO_x electrode after the PEC activity. From Fig. S10, the distribution of all the elements and their quantity was found to be similar to that of TiO₂NR/CdS-NiO_x electrode before PEC activity as shown in Fig. 5. The high-resolution core level XP spectrum of Cd 3d, S 2p and Ni 2p measured from the TiO₂NR/CdS-NiO_x electrode after the PEC activity are shown in Fig. S11. The XPS spectrum of Cd exhibited two peaks arising from Cd 3d_{3/2} and Cd 3d_{5/2} confirming the presence of Cd in (II) state. The deconvoluted XPS spectrum of S 2p show two characteristic peaks arising from S 2p_{3/2} and S 2p_{1/2} ascribed to the sulphide moiety of CdS. The absence of disulfide peak indicates that the CdS is stable under the PEC conditions. The deconvolution of Ni 2p_{3/2} and Ni 2p_{1/2} yielded three peaks corresponding to Ni²⁺, Ni³⁺ and Ni⁰ states indicating the presence of NiO, Ni₂O₃ and metallic Ni. Importantly, the peak ratio of Ni²⁺, Ni³⁺ and Ni⁰ states were found to be approximately equal to those peaks measured for TiO₂NR/CdS-NiO_x electrode before the PEC activity. Overall, the characterization results of TiO₂NR/CdS-NiO_x electrode after the PEC activity indicate that the morphology, crystallographic properties, chemical composition and valency of elements remained unchanged confirming the stability of the proposed p-n junction electrode.

A Mott-Schottky (MS) analysis was utilized to identify the electronic characteristics and flat band potentials of the electrodes. The MS plots were measured at the frequency of 0.5 kHz under the absence of illumination using the following equation:

$$1/C^2 = [2/(N_D e \epsilon_0 \epsilon)] [(V_S - V_{fb}) - (k_B T/e)], \quad (5)$$

where C is the space-charge capacitance, V_S is the applied potential, N_D is the electron carrier density, ϵ_0 is the permittivity of the vacuum, ϵ is the relative permittivity of the semiconductor, e is the elementary charge, T is the temperature (298 K) and k_B is the Boltzmann constant. The plot of $1/C^2$ vs. the applied potential is shown in Fig. 10a-d. Both the TiO₂-NR (Fig. 10a) and TiO₂-NR/CdS (Fig. 10b) electrodes exhibited a positive slope, which indicates an n-type semiconductor with electrons as majority carriers. As expected, the MS plots of NiO_x (Fig. 10c) show a negative slope, thereby confirming a p-type semiconductor. Interestingly, the MS plots of the TiO₂-NR/CdS-NiO_x electrode showed inverted V shaped curve containing both positive and negative slopes (Fig. 10d), which confirmed the formation of a p-n junction [44]. Importantly, the appearance of both negative and positive slopes indicates that the both the n- and p-type materials are in contact with the electrolyte. This could be due to the uneven deposition of NiO_x on the TiO₂-NR/CdS structure. As seen in Fig. 5, the deposition of NiO_x was noticed to be more on the top of the TiO₂-NR/CdS structure compared to its side walls. The crossing point of positive and negative slopes shown in Fig. 10d indicate the capacitance values of the space-charge regions of the junction [45]. Furthermore, Fig. 10d shows the MS plots of TiO₂-NR/CdS-NiO_x p-n junction electrode at different frequencies. From the figure it can be revealed that the capacitance values are dependent on the frequency and the linearity in the graphs indicate the formation of barrier between TiO₂-NR/CdS-NiO_x p-n junction [46]. Further, the evidence of p-n junction formation was obtained by measuring the I-V characteristics of TiO₂-NR/CdS-NiO_x electrode in both forward and reverse bias. The I-V characteristics shown in Fig. 10e produce rectification as in the p-n junction diode confirming the

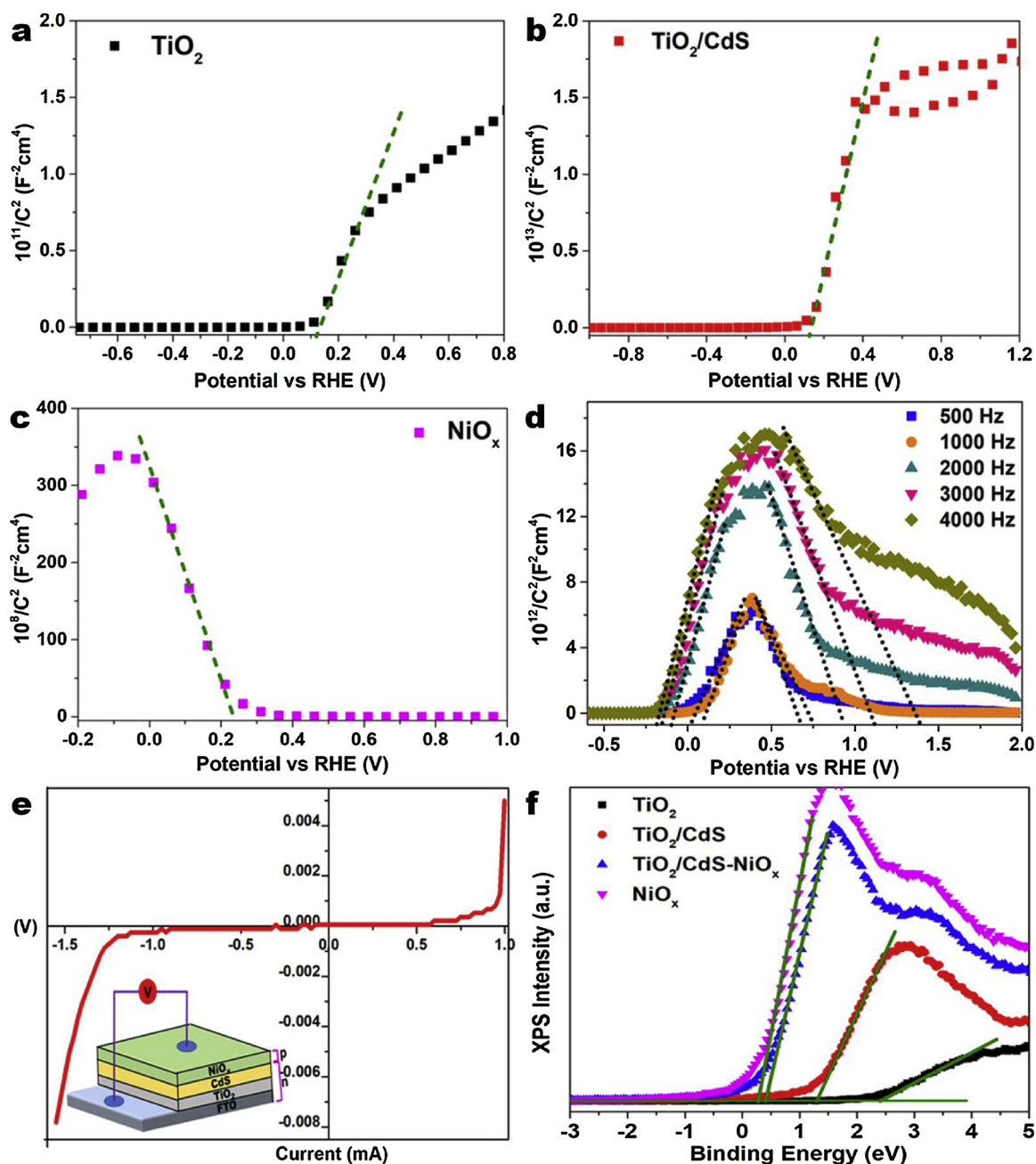


Fig. 10. (a) Mott-Schottky plots of (a) TiO_2 , (b) $\text{TiO}_2\text{-NR/CdS}$, (c) NiO_x , and (d) the $\text{TiO}_2\text{-NR/CdS-NiO}_x$ p-n junction electrode. (d) Mott-Schottky plots of $\text{TiO}_2\text{-NR/CdS-NiO}_x$ p-n junction electrode measured at different frequencies. (e) I-V plots of $\text{TiO}_2\text{-NR/CdS-NiO}_x$ electrode showing typical characteristic of p-n junction. (f) XPS VB spectra of the TiO_2 , $\text{TiO}_2\text{-NR/CdS}$, $\text{TiO}_2\text{-NR/CdS-NiO}_x$ and NiO_x thin films.

formation of $\text{TiO}_2\text{-NR/CdS-NiO}_x$ p-n junction. It can be noted that, when the polarity of bias voltage is altered, the rectification shows an asymmetrical current flow. In the forward bias region, a sudden increase in the current takes place at the “knee” point whereas, during reverse bias sudden rise in current was noticed at the breakdown voltage. Fig. 10f presents the XPS VB edge graphs of NiO_x , $\text{TiO}_2\text{-NRs}$, $\text{TiO}_2\text{-NR/CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ electrodes. The plots show the energy difference between the Fermi level and the VB edge, thereby confirming the type of electrode. That is, the energy differences for TiO_2 and $\text{TiO}_2\text{-NR/CdS}$ was found to be 2.43 and 1.31 eV, respectively, confirming the n-type electrodes, while that of NiO_x was 0.26 eV, indicating its p-type characteristics. However, that of the $\text{TiO}_2\text{-NR/CdS-NiO}_x$ electrode was 0.41 eV, which was mainly due to the NiO_x present on the top of the $\text{TiO}_2\text{-NR/CdS}$ electrode shifting the energy value after the p-n junction formation.

Based on the spectroscopic and electrochemical data, a probable band edge diagram is proposed for the FTO/TiO_2 , FTO/NiO_x , $\text{FTO/TiO}_2/\text{CdS}$ and $\text{TiO}_2\text{-NR/CdS-NiO}_x$ electrode as shown in Fig. 11. Depending on the semiconductor, it can be estimated that the flat-band potential (E_{fb}) should lie close to the VB or CB. Hence, the CB (the TiO_2 NR and $\text{TiO}_2\text{-NR/CdS}$) and VB (NiO_x) edge positions of all the electrodes were estimated by determining the E_{fb} from the intercepts in the MS plots. As can be seen in Fig. 10a, b and c the E_{fb} of the TiO_2 NRs, $\text{TiO}_2\text{-NR/CdS}$ and NiO_x were calculated as 0.116, 0.118 and 0.25 V, respectively. Therefore, the CB edge of TiO_2 and TiO_2/CdS is expected to lie above the E_{fb} having a potential difference of ~ 0.1 eV whereas, the VB edge of NiO_x is expected to be located at ~ 0.1 eV below the E_{fb} . By combining the XPS VB edge and band gap values, the band edge positions of FTO/TiO_2 , FTO/NiO_x , $\text{FTO/TiO}_2/\text{CdS}$ were estimated and plotted as shown in Fig. 11a, b and c, respectively. When n-type TiO_2 -

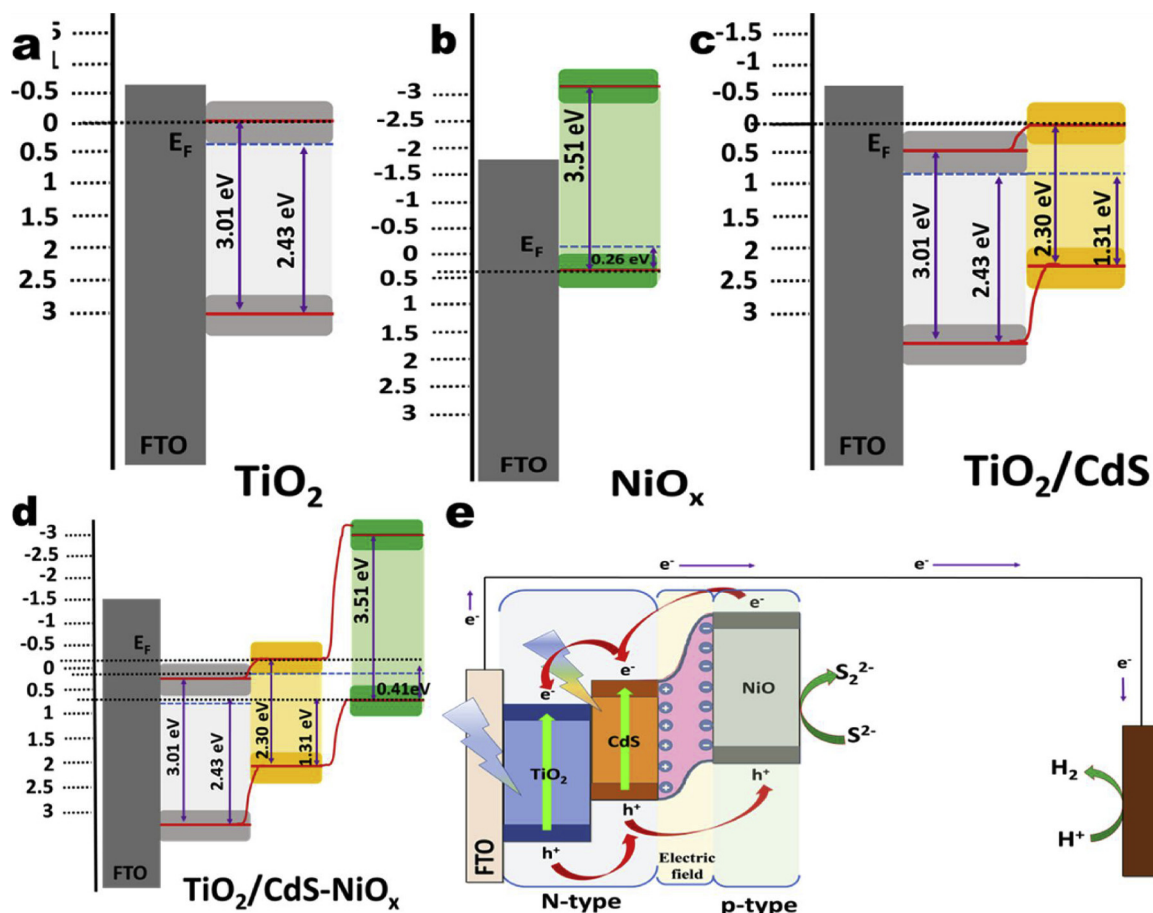


Fig. 11. Band-edge alignment of (a) TiO_2 , (b) NiO_x , (c) $\text{TiO}_2\text{-NR/CdS}$, and (d) the $\text{TiO}_2\text{-NR/CdS-NiO}_x$ p-n junction electrode indicating the VB, CB and Fermi level positions. (e) Schematic diagrams of the p-n junction formation and the PEC water splitting.

NR/CdS and p-type NiO_x were in contact with each other to form a p-n junction, band bending would have occurred at the interface to reach an equal Fermi level, thereby causing an upward shift in the band-edge position of $\text{TiO}_2\text{-NR/CdS}$ and the downward shift of NiO_x , as is evident in Fig. 11d. That is, the intercept of the positive slope observed at 0.767 V shows the typical characteristics of the VB edge of p-type NiO_x on the surface whereas the negative slope intercepted at 0.019 V shows the CB edge of the n-type TiO_2/CdS heterojunction [47]. This suggests that when n-type $\text{TiO}_2\text{-NR/CdS}$ is combined with p-type NiO_x , the band edge positions of the individual materials changed significantly due to band bending and Fermi level alignment, which caused a change in the electron transfer pathway [48]. After the $\text{TiO}_2\text{-NR/CdS-NiO}_x$ p-n junction formation, the electrons near the interface diffused towards the NiO_x region, which left behind positively charged ions in the $\text{TiO}_2\text{-NR/CdS}$ region. Furthermore, the holes near the $\text{TiO}_2\text{-NR/CdS-NiO}_x$ p-n junction interface began to diffuse into the $\text{TiO}_2\text{-NR/CdS}$, thereby leaving ions with a negative charge that led to the formation of a space-charge region at the $\text{TiO}_2\text{-NR/CdS-NiO}_x$ interface. This created an electric field that was able to force photogenerated electrons and holes to transfer in the opposite direction (Fig. 11e). Therefore, the photogenerated electrons and holes in $\text{TiO}_2\text{-NR/CdS-NiO}_x$ were spatially isolated, which led to efficient charge separation. Hence, during the applied bias potential, either an increase or decrease of the band bending could be expected. During anodic polarization, the capacitance of the depletion layer (at the junction) will be minimum favouring the efficient flow of charges. While during the cathodic polarization the capacitance at the junction increases rapidly obstructing the flow of charges [49]. Overall, the characterization results indicate that the enhancement in the PEC properties of $\text{TiO}_2\text{-NR/CdS-NiO}_x$ electrode

could be ascribed to the formation of p-n junction which efficiently separates the photogenerated charges through p-n junction. Additionally, the improved PEC activity was also due to the large drive force for the surface reaction favoured by NiO_x and the low charge transfer resistances across the electrode/electrolyte interface.

4. Conclusion

We successfully fabricated a nanostructured p-n junction electrode consisting of a NiO_x layer on a TiO_2 NR and CdS core-shell structure for efficient photoelectrochemical hydrogen generation. Aligned TiO_2 NRs on FTO and the CdS shell structure on the TiO_2 NRs were synthesised using a hydrothermal method in sequence, and a thin layer of NiO_x was deposited on the top of the $\text{TiO}_2\text{-NR/CdS}$ core-shell structure via RF magnetron sputtering. Characterisation confirmed the formation of the highly crystalline tetragonal, hexagonal and cubic phases of TiO_2 , CdS and NiO_x , respectively. An XPS analysis revealed the presence of NiO_x with various chemical and electronic states, including Ni^{2+} , Ni^{3+} and metallic Ni. An XPS VB edge analysis combined with MS plots confirmed the formation of a $\text{TiO}_2\text{-NR/CdS-NiO}_x$ p-n junction electrode. The $\text{TiO}_2\text{-NR/CdS-NiO}_x$ electrode exhibited a record photocurrent of $\sim 30 \text{ mA cm}^{-2}$ (1 V vs. Ag/AgCl) under full solar spectrum irradiation and an IPCE value of 97% at 500 nm. Furthermore, a superior H_2 -production rate of $\sim 1.3 \text{ mmol h}^{-1}$ under 1 sun illumination with a faradaic efficiency of 95% was recorded. Based on the electrochemical and spectroscopic data, a change in band edge position before and after the p-n junction formation is proposed. The formation of the $\text{TiO}_2\text{-NR/CdS-NiO}_x$ p-n junction facilitated efficient charge separation and diffusion along with the suppression of electron-hole recombination leading to

enhanced PEC hydrogen production activity. The results of the present work demonstrate an efficient strategy for enhancing the PEC water-splitting activity at the $\text{TiO}_2\text{-NR/CdS}$ heterojunction by creating a p–n junction using an NiO_x layer.

Acknowledgements

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Appendix A. Supplementary data

Supplementary material related to this article can be found, in the online version, at doi:<https://doi.org/10.1016/j.apcatb.2019.02.063>.

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